

Chemistry A-Level

Understnading what an excellent decision you have already made 😊

1) Last Lesson	2) Last Month
If chemistry disappeared overnight, what everyday things would you miss the most?	Would you rather memorise facts or solve real-world mysteries?
3) Last Term	4) Last Year
If you could invent one thing to change the world, what would it be?	Why do so many careers value A-Level Chemistry, even outside of science?

Key words:
Chemistry
OCR B
Dream Team 😊



Newent Community School
and Sixth Form Centre

CHEMISTRY A LEVEL

INDUCTION BOOKLET

EXPLORE.
DISCOVER.
SUCCEED.

Your A Level
Chemistry journey
starts here!



THINK BIG
WORK HARD
ACHIEVE MORE

Building knowledge | Developing skills | Achieving excellence



DEEPEN YOUR
KNOWLEDGE



DEVELOP PRACTICAL
SKILLS



THINK
CRITICALLY



ACHIEVE
YOUR POTENTIAL

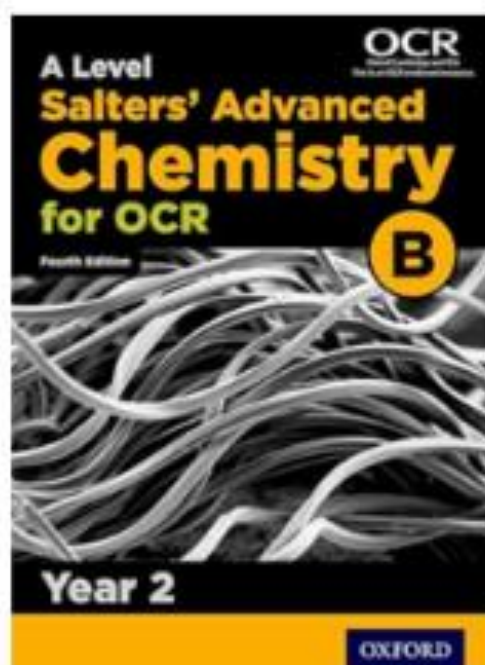
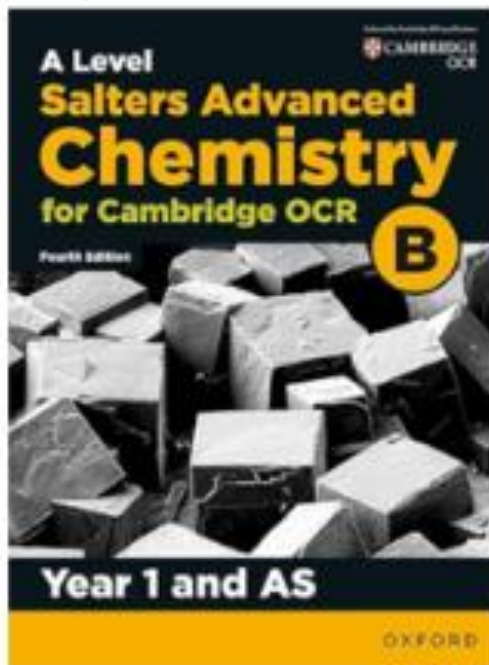


BE PART OF
OUR COMMUNITY

WELCOME TO YOUR
**A LEVEL
CHEMISTRY**
JOURNEY!

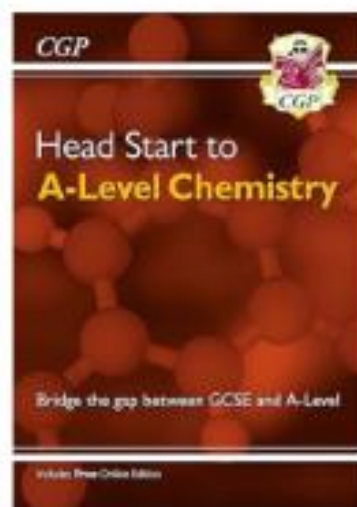
Recommended Textbooks

- OCR A Level Chemistry B (Salters) Student Book (can borrow with a deposit or buy your own)



Bridging GCSE to A-Level:

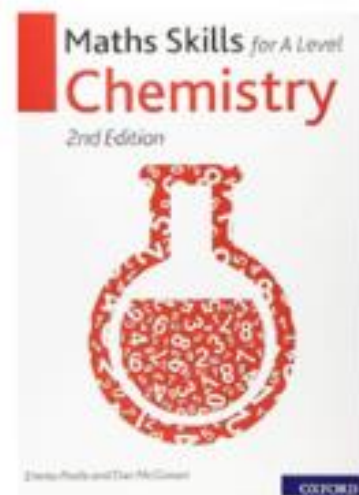
- CGP Head Start to A-Level Chemistry Chemistry



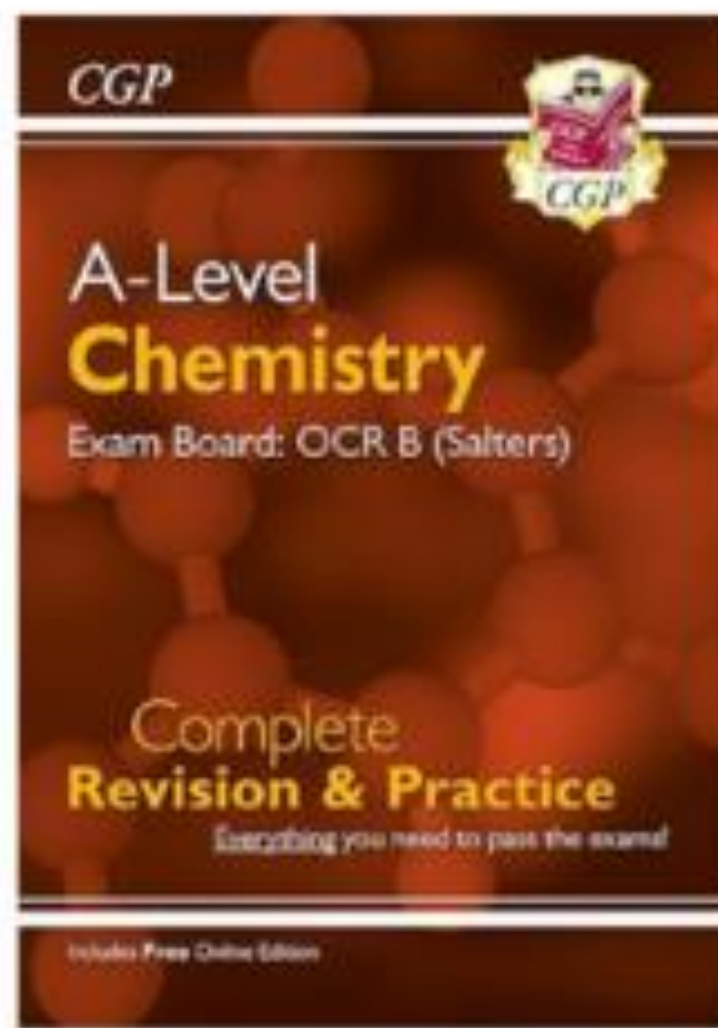
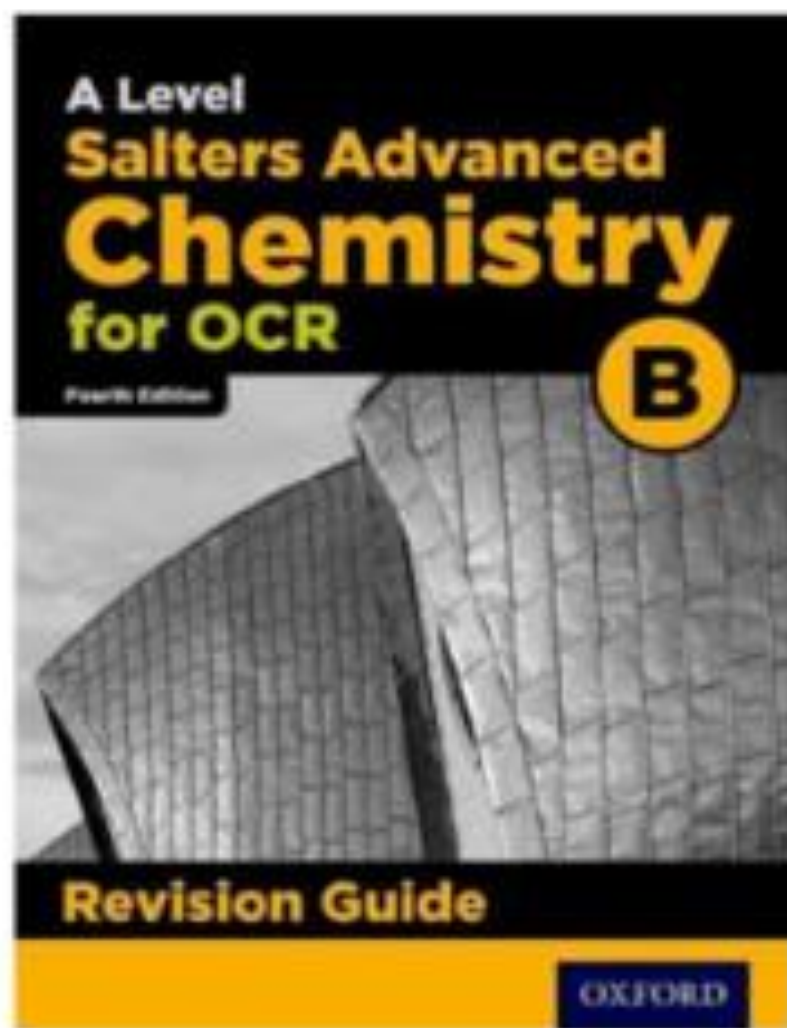
Mathematics Support:

CGP Maths for A-Level

(deposit/purchase)



Recommended Revision:



Study Expectations

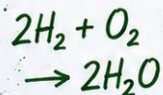
Successful students review notes weekly, complete homework promptly, practise calculations regularly and spend approximately five hours each week studying independently outside lessons.

Summer Work

- Complete before September:
 - 1) Summer revision notes (from the CGP book or other GCSE revision resources).
 - 2) Your completed GCSE exam questions **marked using the mark schemes.**

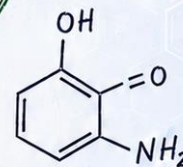


NEWENT COMMUNITY SCHOOL AND SIXTH FORM CENTRE



Chemistry A LEVEL Summer Work

Think
React
Succeed!



5 B Boron 10.81	6 C Carbon 12.01	7 N Nitrogen 14.01	8 O Oxygen 16.00
13 Al Aluminium 26.98	14 Si Silicon 28.09	15 P Phosphorus 30.97	16 S Sulfur 32.06
31 Ga Gallium 72.63	32 Ge Germanium 72.63	33 As Arsenic 74.92	34 Se Selenium 78.96

The best
reactions ☆
start with
curiosity.



Curious
Creative
Confident



2.2 Determination of the molar mass of an acid

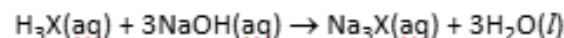
LEARNER

Introduction

In this Practical Activity, you are supplied a solid sample of an unknown hydrated acid. By preparing a solution of the sample and titrating against a standard solution of sodium hydroxide, you will determine the amount of moles of the acid in the sample and hence the molar mass of the acid. To accurately find the amount of acid, you must make the hydrated acid up to an accurately known volume, in the same way as making up a standard solution.

The indicator you will be using in this activity is phenolphthalein, which is colourless in acid and pink in alkali. At the end point of the titration, the colour is **faint pink** (between **colourless** and **pink**).

The equation below represents the reaction that you will carry out:



Aims and Skills

- to accurately prepare a solution of a hydrated acid, $\text{H}_3\text{X} \cdot n\text{H}_2\text{O}$
- to determine the molar mass of the acid by titration against a standard alkali solution
- *to accurately measure mass and volume of liquid*
- *to use laboratory apparatus for titration, using burette and pipette*
- *to accurately make up a standard solution using a volumetric flask*
- *to use acid-base indicators in titration of a strong acid with a weak alkali*
- *to safely handle solids and liquids*